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③ Magnetic Q.N.

Atomic orbitals can have the same shape but different orientations around the nucleus.
read pg 108

Symbolized by the letter m

values allowed are including 0
 $-l$ to $+l$.

exception is s because it is spherical it has only one possible orientation.

n	l	m_l	
1	0 (s)	0	one orientation.
2	0, 1	-1, 0, +1	3 orientations.
3	0, 1, 2	-2, -1, 0, +1, +2	5 orientations.

④ Spin Quantum #, read pg 110

This is the magnetic property of the e^-

A single orbital can hold a max. of 2 electrons, which must have opposite spin state.

Each atomic orbital is designed by
1- principal Q.N followed by a letter of the sublevel.

HW all sec. review including C.T.