

## Isotopes

\* Define the following:

1. Nuclide: a general term for a specific isotope of an element.
2. mass number: # of protons + # of neutrons.
3. atomic number: # of protons

Isotopes usually identified by their mass number

4. Average atomic mass: weighted average of atomic masses of the naturally occurring isotopes of an element

\* Relating Mass to Number of atoms.

Define the following:

1. The Mole: SI unit for amount of substance, it is the amount of a substance that contains as many as Avogadro's number.
2. Avogadro's number:  $6.022 \times 10^{23}$  particles
3. Molar Mass: Mass of one mol of pure

Continue →