

① Principal Quantum #

read pg 107
 values of n allowed are the integers 1, 2, 3, ...
 n = main energy level (integers 1, 2, 3, ...)
 as n increase means its farther away from the nucleus.
 n^2 = the total # of orbitals exist in a given level.

② Angular momentum Quantum # Symbolized by l
 indicates the shape of the orbital
 define read pg 107

* values of l allowed are 0 and any integer less than or equal to $n-1$

$$(0) \rightarrow (n-1)$$

each orbital is assigned a letter.

we are also going to call them sublevels.
 $n=1$ has one sublevel s
 $n=2$ has 2 sublevels s, p
 $n=3$ has 3 sublevels s, p, d
 $n=4$ has 4 sublevels s, p, d, f

$l \rightarrow n-1$	n	$l \rightarrow (n-1)$
0, 1, 2	1	0 (s)
	2	0, 1
	3	0, 1, 2
	4	0, 1, 2, 3

Balloons