

Explain how if there is a gradual increase in atomic radii down a group, there will also be a gradual increase in the ionic radii.

- As the atom's radius gets smaller, the electrons get closer together, causing a stronger attraction between the protons and the electrons.

In class Explanation: As they are in atoms, the outer electrons in both cations & anions are in high energy levels as one reads down a group. Therefore, just there is a gradual increase of atomic radii down a group, there is gradual increase ionic radii.

Is the electronegativity related to the size of the atom? Explain in terms of the attraction force between protons & electrons?

Since electrons have technically almost no mass, if an electron is added it wouldn't really make a difference. However, if an electron has mass it has size so I'm not sure.