Chapter 5 Test.
Continue…

5. Explain the basis of each trend.
   a. The first ionization energies of main group elements across a period.

   b. First ionization energies of main group elements down a group.

6. The two ions K^+ and C^{2+} each have two electrons surrounding the nucleus. Which would you expect to have the smaller radius? Why?

7. Describe some differences between s-block metals and d-block metals.

8. Why do the halogens readily form one –ve ions?

9. The electron configuration of argon differs from those of chlorine and potassium by one electron. Compare the reactivity of these two elements.